



Chemistry Content Standards:

The kinetic molecular theory describes the motion of atoms and molecules and explains the properties of gases.

Standard 4a: Students know the random motion of molecules and their collisions with a surface create the observable pressure on that surface.

Standard 4b: Students know the random motion of molecules explains the diffusion of gases.

Standard 4c: Students know how to apply the gas laws to relations between the pressure, temperature, and volume of any amount of an ideal gas or any mixture of ideal gases.

- 13.1.1 Describe the assumptions of the kinetic theory as it applies to gases
- 13.1.2 Interpret gas pressure in terms of kinetic energy
- 13.1.3 Define the relationship between Kelvin temperature and average kinetic energy
- 13.1.4

Vocabulary: Chapter 13 and 14

Lecture: Gases 13.1 PPT

Pressure and Temperature Page 52 in NB

Chemistry

Principles of flight, intro to Chapter 6.1

Dalton's Law

Charles Law

Describing Gases, V, T, P

Classwork : Work sheet 13.1 The Nature of Gases
WB 13.1
Vocabulary 13

Homework: Page 389, 3-6